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 Worksheet- Calculations involving Specific Heat

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### JEFFERSON SHANIA

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#### Chemistry Practice Problems: Heat and Specific Heat

Worksheet 16 — Calorimetry Calorimetry is the experimental measurement of heat (q) produced in chemical and physical processes. Heat can not be measured directly, but temperature changes can be measured. The factor that links these two is heat capacity. Heat capacity, C, is defined as the heat required to raise the temperature of a

Worksheet- Calculations involving Specific Heat. 1. For  $q = m \cdot c \cdot \Delta T$  : identify each variables by name & the units associated with it. q = amount of heat (J) m = mass (grams) c = specific heat (J/g°C)  $\Delta T$  = change in temperature (°C) 2. Heat is not the same as temperature, yet they are related.

#### Specific Heat Show all calculations for credit. Chem ...

of energy. Calculate the specific heat of the metal and determine its identity. Specific Heat Name \_\_\_\_ Chem Worksheet 16-1 Useful Equations  $q = mc \Delta T$   $c = 5/9(T_f - 32)$   $T = T_{\text{final}} - T_{\text{initial}}$   $T_K = T_c + 273$   $1 \text{ kg} = 1000 \text{ g}$   $1 \text{ kcal} = 1000 \text{ cal}$   $1 \text{ cal} = 4.184 \text{ J}$

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#### Specific Heat Worksheet - Broadneck High School

These example problems show you how to solve for heat, specific heat, and the change in temperature using  $q = m \cdot C \cdot \Delta T$ . Includes metric unit conversions and conversions between calories and joules.

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Honors Chemistry Worksheet - Specific Heat. Recognize that when two systems at different temperatures meet, there will be a net transfer of heat (energy) from the system of greater heat intensity to the system of lower heat intensity. Summary - Heat flows from source to sink, in other words from hot to cold until thermal equilibrium is obtained.

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If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature. (Note: 1 ml of water = 1 gram) 5. A total of 54.0 kilojoules of heat are absorbed as 58.3 g of lead is heated from 12.0°C to 42.0°C.

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