

# Chapter 7 Chemical Formulas And Chemical Compounds

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Chemistry Chapter 7 Chemical Formulas and Chemical ... Chapter  
 7 Chemical Formulas And CHAPTER 7 REVIEW Chemical Formulas  
 and Chemical Compounds SECTION 3 SHORT ANSWER Answer  
 the following questions in the space provided. 1. Label each of  
 the following statements as True or False: True a. If the formula  
 mass of one molecule is  $x$  amu, the molar mass is  $x$  g/mol. False  
 b. Samples of equal numbers of moles of two different  
 chemicals 7 Chemical Formulas and Chemical  
 Compounds chemical formula reveals the number of atoms of  
 each element contained in a single molecule of the compound,  
 as shown below for the hydrocarbon octane. (Hydrocarbons are  
 molecular compounds composed solely of carbon and hydrogen.)  
 CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS 203 SECTION  
 7-1 OBJECTIVES Explain the significance of a chemical  
 ... CHAPTER 7 Chemical Formulas and Chemical  
 Compounds Chapter 7: Chemical Formulas and Chemical  
 Compounds Section 7-1: Chemical Names and Formulas 7-1-1  
 Explain the significance of a chemical formula. The subscript  
 indicates that there are 8 carbon atoms in the molecule of  
 octane. The subscript indicates that there are 18 hydrogen atoms  
 in the molecule of octane. Subscript 2 refers to 2 aluminum  
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 Formulas Flashcards | Quizlet Chapter 7 Significance of a  
 Chemical Formula • A chemical formula indicates the relative  
 number of atoms of each kind in a chemical compound. • For a  
 molecular compound, the chemical formula reveals the number  
 of atoms of each element contained in a single molecule of the  
 compound. • example: octane —  $C_8H_{18}$  The subscript after the  
 C Chemical Formulas and Chemical Chapter 7 Compounds Chapter  
 7 Significance of a Chemical Formula • A chemical formula  
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 single molecule of the compound. • example: octane —  $C_8H_{18}$   
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 Compounds Test B ... Chapter 7: Chemical Formulas and Chemical  
 Compounds Section 2: Oxidation Numbers Oxidation numbers An  
 oxidation number, also called oxidation state, indicates the  
 general distribution of electrons among the bonded atoms in a  
 molecular compound or polyatomic ion. Unlike ionic charges,  
 oxidation numbers do not have an exact physical  
 meaning. Chapter 7: Chemical Formulas and Chemical  
 Compounds Chapter 7 - Chemical Formulas & Chemical  
 Compounds Chapter 7 is the final chapter of the first semester. It  
 serves as a summary of all course content discussed so far, as  
 the focus of this chapter. ... Chapter 7 - Chemical Formulas &  
 Chemical Compounds - yazvac 7. What is the empirical formula for  
 lactic acid, which has a molecular formula of  $C_3H_5O_3$ ? a. b.  
 $CHO_{232}$  c.  $CHO_{363}$  d.  $CHO_8$ . NaCl is a. a molecular formula  
 only. b. an empirical formula only. c. both an empirical formula  
 and a chemical formula. d. both a molecular formula and an  
 empirical formula. 9. brearleyhigh.kenilworthschools.com Start  
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 — Chemical Formulas and Nomenclature page 1 WORKSHEET 1:

Determination of oxidation number or valence number Rules to apply: 1. a. The net charges on all molecules is zero; therefore, the sum of the positive charges equals the sum of the negative charges 2. a. WORKSHEET 1: Determination of oxidation number or valence ... Chapter 7 - Chemical Formulas and Chemical Compounds 7-1 Chemical Names and Formulas I. Significance of a Chemical Formula A. Molecular formulas 1. Number of atoms of each element in one molecule of a compound C<sub>2</sub>H<sub>6</sub> = ethane (2 carbon atoms, 6 hydrogen atoms) B. Ionic Compounds 1. Chapter 7 - Chemical Formulas and Chemical Compounds CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO<sub>3</sub> a. copper(II) carbonate Na<sub>2</sub>SO<sub>3</sub> b. sodium sulfite (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> c. ammonium phosphate SnS<sub>2</sub> d. tin(IV) sulfide HNO<sub>2</sub> e. nitrous acid 2.7 Chemical Formulas and Chemical Compounds CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Assign the oxidation number to the specified element in each of the following examples: a. S in H<sub>2</sub>SO<sub>3</sub> b. S in MgSO<sub>4</sub> c. S in K<sub>2</sub>S d. Cu in Cu<sub>2</sub>S e. Cr in Na<sub>2</sub>CrO<sub>4</sub> f. N in HNO<sub>3</sub> g. C in (HCO<sub>3</sub>)<sup>-</sup> h. N in (NH<sub>4</sub>)<sup>+</sup> 2. a. Chemical Formulas and Chemical Compounds - MAFIADOC.COM CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In a Stock system name such as iron(III) sulfate, the Roman numeral tells us (a) how many atoms of Fe are in one formula unit. (b) how many sulfate ions can be attached to the iron atom. (c) the charge on ... 7 Chemical Formulas and Chemical Compounds Test and improve your knowledge of Holt McDougal Modern Chemistry Chapter 7: Chemical Formulas and Chemical Compounds with fun multiple choice exams you can take online with Study.com

chemical formula reveals the number of atoms of each element contained in a single molecule of the compound, as shown below for the hydrocarbon octane. (Hydrocarbons are molecular compounds composed solely of carbon and hydrogen.)

CHEMICAL FORMULAS AND CHEMICAL COMPOUNDS 203 SECTION 7-1 OBJECTIVES Explain the significance of a chemical ...

Chapter 7 - Chemical Formulas and Chemical Compounds 7-1 Chemical Names and Formulas I. Significance of a Chemical Formula A. Molecular formulas 1. Number of atoms of each element in one molecule of a compound C<sub>2</sub>H<sub>6</sub> = ethane (2 carbon atoms, 6 hydrogen atoms) B. Ionic Compounds 1.

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Assign the oxidation number to the specified element in each of the following examples: a. S in H<sub>2</sub>SO<sub>3</sub> b. S in MgSO<sub>4</sub> c. S in K<sub>2</sub>S d. Cu in Cu<sub>2</sub>S e. Cr in Na<sub>2</sub>CrO<sub>4</sub> f. N in HNO<sub>3</sub> g. C in (HCO<sub>3</sub>)<sup>-</sup> h. N in (NH<sub>4</sub>)<sup>+</sup> 2. a.

**Chapter 7 - Chemical Formulas and Chemical Compounds**

Chapter 7 Significance of a Chemical Formula •A chemical formula indicates the relative number of atoms of each kind in a chemical compound. •For a molecular compound, the chemical formula reveals the number of atoms of each element contained in a single molecule of the compound. •example: octane — C<sub>8</sub>H<sub>18</sub> The subscript after the C

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Chemistry Chapter 7 Worksheets —Chemical Formulas and Nomenclature page 1 WORKSHEET 1: Determination of oxidation number or valence number Rules to apply: 1. a. The net charges on all molecules is zero; therefore, the sum of the positive charges equals the sum of the negative charges 2. a. [chemistry chapter 7 chemical formulas Flashcards and Study ...](#) Learn chemistry chapter 7 chemical formulas with free interactive flashcards. Choose from 500 different sets of chemistry chapter 7 chemical formulas flashcards on Quizlet.

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*7 Chemical Formulas and Chemical Compounds*

7. What is the empirical formula for lactic acid, which has a molecular formula of (C<sub>3</sub>H<sub>5</sub>O<sub>3</sub>)<sub>n</sub>? a. b. CHO<sub>2</sub> c. CHO<sub>3</sub> d. CHO<sub>8</sub> 8. NaCl is a. a molecular formula only. b. an empirical formula only. c. both an empirical formula and a chemical formula. d. both a molecular formula and an empirical formula. 9. *Chemical Formulas and Chemical Compounds - MAFIADOC.COM* Start studying Chapter 7: Chemical Formulas. Learn vocabulary, terms, and more with flashcards, games, and other study tools.

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formula reveals the number of atoms of each element contained in a single molecule of the compound. •example: octane — C<sub>8</sub>H<sub>18</sub> The subscript after the C

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