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 REVIEW Stoichiometry SECTION 3  
 PROBLEMS Write the answer on the line  
 to the left. Show all your work in the  
 space provided. 1. 88% The actual yield  
 of a reaction is 22 g and the theoretical  
 yield is 25 g. Calculate the percentage  
 yield. 2. 6.0 mol of N<sub>2</sub> are mixed with  
 12.0 mol of H<sub>2</sub> are mixed with  
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 6.022 × 10<sup>23</sup> particles (atoms, molecules)  
 ( 2. 1 mole = 6.022 × 10<sup>23</sup> particles (

1 mole = molar mass (1 mole = 22.4 L (3.  
 Calculate (the ... Answer Key  
 Mole/Stoichiometry. Test. Review stoichio  
 metry (which you studied in Chapter 3)  
 deals with the mass relationships of  
 elements in compounds. Reaction  
 stoichiometry involves the mass  
 relationships between reactants and  
 products in a chemical reaction.  
 Reaction stoichiometry is the subject of  
 this chapter and it is based on CHAPTER  
 9 Stoichiometry - Riverside Local  
 Schools Remember it is a MC test, use  
 the answers ... 3 contains 9 × 10<sup>23</sup>  
 oxygen atoms. III. A 200 g sample of  
 CaCO<sub>3</sub> contains 2 moles of CaCO<sub>3</sub> a. I  
 only b. II only c. III only d. I and III only e.  
 I, II, and III Practice Test Ch 3  
 Stoichiometry Name \_\_\_\_\_ Per \_\_\_\_\_ 2MnO  
 2 + 4KOH + O<sub>2</sub> + Cl<sub>2</sub> → 2KMnO<sub>4</sub> +

2KCl + 2 H<sub>2</sub>O. For the reaction above, there ...Practice Test Ch 3 Stoichiometry Name PerChapter 9 focuses on reaction stoichiometry: using a balanced chemical equation to calculate the number of grams, moles, or particles of reactants/products involved in a chemical reaction. Students...Chapter 9 - Stoichiometry - yazvac - Google SitesChapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: 2H<sub>2</sub>O<sub>2</sub> (aq) → 2H<sub>2</sub>O(l) + O<sub>2</sub> (g) This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

CHAPTER 9 REVIEW Stoichiometry

SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H

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Chapter 9 Test Stoichiometry Answers

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on Modern chemistry chapter 9 review stoichiometry answers

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Chapter 9: Standard Review Worksheet

1. Answers will vary. An example is included below:  $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l})$

+  $\text{O}_2(\text{g})$  This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

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relationships of elements in compounds.



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9. If the percentage yield for a chemical  
reaction is 80.0%, and the theoretical  
yield is 100 grams, what is the actual  
yield of product? ... Chapter 9

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### **Test Answers**

Remember it is a MC test, use the  
answers ... 3 contains  $9 \times 10^{23}$  oxygen  
atoms. III. A 200 g sample of  $\text{CaCO}_3$   
contains 2 moles of  $\text{CaCO}_3$  a. I only b. II  
only c. III only d. I and III only e. I, II, and  
III Practice Test Ch 3 Stoichiometry  
Name \_\_\_\_\_ Per \_\_\_\_\_  
 $2\text{MnO}_2 + 4\text{KOH} + \text{O}_2 + \text{Cl}_2 \rightarrow 2\text{KMnO}_4 + 2\text{KCl} + 2\text{H}_2\text{O}$  9.  
For the reaction above, there ...

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Answer Key

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$6.022 \times 10^{23}$  particles ((atoms,(molecules)(

( 2. 1mole(= $6.022 \times 10^{23}$  particles((

1mole=molar(mass(1mole=22.4L(3.

Calculate(the ...