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 Some of the alpha particles were deflected (“deflection” is an observation, “repulsion” is not). (c)
 Rutherford included several features in the experiment due to the characteristics of the alpha
 particles.ATOMS: ATOMIC STRUCTURE QUESTIONS AND ANSWERSAtomic Structure. Revision
 Questions. The best way to remember the information in this chapter is to get a pen and paper and
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 electrons. To answer the question, you need to know the electron configuration of the atoms to see
 which one has 7 unbonded electrons, like chlorine. Fluorine, element number 9, has 2 electrons in
 the s sublevel (K shell).Atomic Structure Chemistry Quiz - thoughtco.comWhen alpha particles are
 used to bombard gold foil, most of the alpha particles pass through undeflected. This result
 indicates that most of the volume of a gold atom consists of _____.Atomic Structure Multiple Choice
 Questions(b) An atom of oxygen has approximately 16 times the mass of an atom of hydrogen. In
 the formula $\lambda = h/mv$, h is constant while the conditions of problem make v, also constant. This
 means that λ and m are variables and λ varies inversely with m.Solved Questions of Atomic
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 & Fundamentals of Matter - Study.comQ. How many electrons are in an atom of carbon (C)?
 (Remember APE! Atomic Number = # of Protons and # of Electrons)Atomic Structure | Atoms &
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 first time. What was the subject called so far? Answer: Natural philosophy. 3) Which Greek
 philosopher suggested that the material world was composed of four basic elements – air,
 water,...118 Physics Quiz Questions Answers - Learn about Physics ...Chemistry Quiz Questions -
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 neutrons. neutrons and electrons. electrons and protons. The charge and mass number of an
 electron are: charge = +1, Mass number = 0. charge = -1, Mass number = 0. charge = 0, Mass
 number = 1.Atomic Structure ReviewSample exam questions - Atomic structure and the periodic
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 Q. How many electrons are in an atom of carbon (C)? (Remember APE! Atomic Number = # of
 Protons and # of Electrons)
 Atomic Structure Review. The nucleus of most atoms is made up of: protons and electrons. protons
 and neutrons. neutrons and electrons. electrons and protons. The charge and mass number of an
 electron are: charge = +1, Mass number = 0. charge = -1, Mass number = 0. charge = 0, Mass
 number = 1.
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1.1 Exercise 1 - atomic symbols 1.1 Exercise 2 - ram, rmm and mass spectra 1.1 Exercise 3 - electronic configuration 1.1 Exercise 4 - ionisation energies Answers to 1.1 Exercises

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Multiple choice questions - Sample exam questions - Atomic ...

(ii)The nucleus is positively charged. Some of the alpha particles were deflected ("deflection" is an observation, "repulsion" is not). (c) Rutherford included several features in the experiment due to the characteristics of the alpha particles.

ATOMS: ATOMIC STRUCTURE QUESTIONS AND ANSWERS

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Quiz. Electrochemical Reactions Chemistry Questions. Periodic Classification. 1. What is the optimum temperature in the manufacture of ammonia? 2. Name the scientist who described about reactions that go forward and backward.

Atomic Structure Multiple Choice Questions

(b) An atom of oxygen has approximately 16 times the mass of an atom of hydrogen. In the formula $\lambda = h/mv$, h is constant while the conditions of problem make v , also constant. This means that λ and m are variables and λ varies inversely with m .